**WODSS SCIENCE** Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

SCH 4CI Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**3.3 Fractional Distillation and Cracking**

What is petroleum? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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Small hydrocarbon molecules exist as \_\_\_\_\_\_\_\_\_\_\_\_\_. Some examples are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Larger hydrocarbon molecules are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. Some examples are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

The heaviest hydrocarbons are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Why are the most valuable hydrocarbons in petroleum the ones with 5 to 12 carbon atoms?

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Different hydrocarbon molecules have different boiling points. Which ones have the lowest boiling points? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Which ones have the highest boiling points? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Hydrocarbons are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ (polar or non-polar) molecules, and so the intermolecular forces between molecules are weak London dispersion forces. This force gets stronger as the length of the hydrocarbon molecules \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. So, bigger molecules have higher boiling points because more energy is needed to pull the molecules apart to change them into a gas.

**Fractional Distillation**

The separation of components of petroleum by distillation, using differences in boiling point is called \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. Different sized molecules are separated into groups called \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. The lighter fractions boil at \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ temperatures and the heavier fractions boil at \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ temperatures.

*Put the steps of fractional distillation below in order by numbering them 1 🡪 5.*

*Step #*

\_\_\_\_\_ Larger molecules condense into liquid at the lower warmer sections.

\_\_\_\_\_ The liquids from each fraction are collected on a tray.  
\_\_\_\_\_ Entire mix of hydrocarbons is heated to very high temperatures so all the hydrocarbons evaporate.

\_\_\_\_\_ Smaller molecules condense into liquid at the higher cooler sections.

\_\_\_\_\_ Hot gasses rise in a tall fractionation tower, which is warm at the bottom, and cooler at the top.

**Cracking**

The process in which large straight-chain hydrocarbon molecules are converted into shorter branched-chain hydrocarbon molecules is called \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

How is a hydrocarbon cracked? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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What else can cracking do to hydrocarbons? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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Answer the following questions on the back or on a separate piece of paper: p. 196 # 1 -6.