

The Atom – Review

1. Use your notes to complete the following questions.

Sub atomic particle	Relative mass	Charge	Location in atom
proton			
electron			
neutron			

2. How do you calculate the number of protons, electrons and neutrons? Use the words “atomic number” and / or “mass number” to describe how to calculate the number of each.

of protons = _____

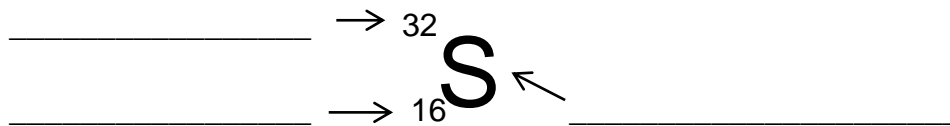
of electrons = _____

of neutrons = _____

3. The atomic number is especially important because it tells you what type of element you have. Use your periodic table to complete the chart below. Write down the element name and its symbol with the following atomic numbers.

Atomic number	Element name	Element symbol
6		
22		
80		

4. Standard atomic Notation – review



5. Complete the following chart. Use your periodic table, and your answers to question 2 above.

Element name	Symbol	Standard atomic notation	Atomic number	Mass number	number of protons	number of electrons	number of neutrons
Helium							
	Be						
		22					
					53		
						92	

6. What is the difference between observation and inference?

A nurse examining a patient with a high fever

Observation: _____

Inference: _____

7. What contribution did the following people make to the atomic theory? Include a picture along with the description.

- a. Dalton b. Thomson c. Chadwick d. Rutherford e. Bohr