+ an element - 6.022 x 10^23 atoms of that element.
: CaCl2 (I.C) - 6.022 x 10^23 F.U.

Avogadro's Constant and the Mole

istry, we don't talk about a certain number of atoms because atoms are too small to count. We'd be talking in umbers if we were counting atoms. Instead, we use a called a mole.

ole = 6.022×10^{23} particles (atoms, ions, molecules or ıula units)

= 602 200 000 000 000 000 000 000 particles!!!

H_2O (m.c) - 6.022×10^{23} molecules of water.

ıber 6.022×10^{23} is called **Avogadro's constant** (N_A).

o realized that 6.022×10^{23} atoms of any element have a grams, that is equal to the numerical value of the 's atomic mass.

e – one atom of carbon has a mass of 12.011 μ or amu,
mole of carbon atoms has a mass of 12.011 g.
m of iron has a mass of 55.845 μ, and one mole of iron as a mass of 55.845 g

use Avogadro's constant to calculate the number of moles stance, or the number of particles in a substance. We use wing symbols in our calculations: $n = \frac{N}{N_A}$

Quantity	Unit
# of moles	mol
# of particles	atoms, ions, formula units or molecules
Avogadro's number (6.022×10^{23})	particles / mol

w that 1 mole = 6.022×10^{23} particles. List the possible ion factors from this equation:

1 mol (2) 6.022×10^{23} ...