# Acids, Acid Salts and Bases Nomenclature

| Write the chemical formula | | Write the chemical name | |
|---|---|---|---|
| 1 | hydrochloric acid $HC\ell_{(aq)}$ | 26 | $HNO_{3(aq)}^{(classical)}$ nitric acid |
| 2 | calcium peroxide $CaO_2$ | 27 | $HC\ell O_{3(aq)}$ chloric acid |
| 3 | ammonium sulphide $(NH_4)_2S$ | 28 | $Ca(OH)_{2(aq)}$ calcium hydroxide |
| 4 | calcium oxide $CaO$ | 29 | $H_3PO_{4(aq)}^{(classical)}$ phosphoric acid |
| 5 | hydroiodic acid $HI_{(aq)}$ | 30 | $HNO_{2(aq)}^{(classical)}$ nitrous acid |
| 6 | aluminium peroxide $AlO_2$ | 31 | $LiHCO_3$ lithium hydrogen carbonate |
| 7 | ammonium bromide $NH_4Br$ | 32 | $Ba(HSO_5)_2$ barium hydrogen persulfate |
| 8 | aluminum oxide $Al_2O_3$ | 33 | $H_3PO_{2(aq)}^{(classical)}$ hypophosphorous acid |
| 9 | ammonium hydroxide $NH_4OH$ | 34 | $Cu(HCO_3)_2$ copper (II) hydrogen carbonate |
| 10 | barium hydroxide $Ba(OH)_2$ | 35 | $H_2SO_{5(aq)}^{(classical)}$ persulfuric acid |
| 11 | hydrosulfuric acid $H_2S_{(aq)}$ | 36 | $NH_4OH$ ammonium hydroxide |
| 12 | hydrogen peroxide $H_2O_2$ | 37 | $NaHSO_3$ sodium hydrogen sulfite |
| 13 | ammonium iodide $NH_4I$ | 38 | $HNO_{2(aq)}$ aqueous hydrogen nitrite Nitrous acid |
| 14 | hydrogen sulfide gas $H_2S$ | 39 | $Ti(OH)_{4(aq)}$ aqueous titanium (IV) hydroxide |
| 15 | lithium hydroxide $LiOH$ | 40 | $HBrO_{4(aq)}$ aqueous hydrogen perbromate perbromic acid |
| 16 | hydrogen chloride gas $HC\ell$ | 41 | $Fe(OH)_3$ iron (III) hydroxide |
| 17 | magnesium hydroxide $Mg(OH)_2$ | 42 | $NaH_2PO_4$ sodium dihydrogen phosphate |
| 18 | potassium peroxide $K_2O_2$ | 43 | $HNO_{3(aq)}$ aqueous hydrogen nitrate Nitric acid |
| 19 | hydrofluoric acid $HF_{(aq)}$ | 44 | $NH_4HCO_2$ ammonium hydrogen carbonite |
| 20 | calcium hydroxide $Ca(OH)_2$ | 45 | $MgHPO_4$ magnesium hydrogen phosphate |
| 21 | cesium oxide $Cs_2O$ | 46 | $HBrO_{3(aq)}$ aqueous hydrogen bromate bromic acid |
| 22 | hydrobromic acid $HBr_{(aq)}$ | 47 | $H_2SO_{3(aq)}$ aqueous hydrogen sulfite sulfurous acid |
| 23 | aluminum hydroxide $Al(OH)_3$ | 48 | $CH_3COOH_{(aq)}$ acetic acid |
| 24 | hydrogen oxide $H_2O$ | 49 | $HIO_{3(aq)}^{(classical)}$ iodic acid |
| 25 | hydrogen iodide gas $HI$ | 50 | $NiHCO_3$ nickel (I) hydrogen carbonate |