

Pre-Lab for Next Class**Read Lab on page 282-283**

1. Title: % Composition By Mass and Empirical Formula Lab
2. Question: What is the percentage composition by mass of magnesium oxide? Using your experimental data find the empirical formula of the compound?
3. Write a balanced Chemical Equation
4. Pre-lab questions #1-3
5. Observation Table with Creative and Descriptive Title
6. Questions/Analysis:
 - a. Calculate the % composition of magnesium oxide (from the chemical formula).
 - b. Calculate the % composition of magnesium oxide (from you lab data).
 - c. Compare the values you got from question and 2. Why might they be different?
 - d. Determine the empirical formula (from your lab data).
 - e. If some of the magnesium oxide had escaped from the crucible, would your percentage composition calculation of magnesium be too high or too low? Explain.
 - f. If the magnesium had reacted with some other component in the air, would your percentage composition calculation of magnesium be too high or too low? Explain.